

S.#	Question	Option (A)	Option (B)	Option (C)	Option (D)	Correct Answer
1	The term which shows the quantitative relationship between the products and reactants is called as _____?	Percentage yield	Stoichiometry	Limiting reactant	All of these	B
2	In the reaction of burning of an organic compound in the presence of oxygen, which of the following is a reagent present in excess?	Organic compound	Oxygen	Fuel	All of these	B
3	If uncertainty in momentum of electron is zero, the uncertainty in its position would be _____?	Less than zero	more than zero	one	infinite	D
4	In the discharge tube positive rays move in which direction?	Towards anode	Towards cathode	move randomly	Move towards vacuum in the tube	B
5	The unit _____ is commonly used by meteorologist	Bar	centibar	millibar	kilobar	C
6	If pressure and volume of a gas are variable while temperature remains constant, this belongs to	Charles's law	Boyle's law	Avogadro's law	Pascal's law	B
7	If a graph is plotted between pressure on x-axis and volume on y-axis for Boyle's law verification, the curve obtained is called	pseudothorn	isotherm	biotherm	all of these	B
8	All the halogens are _____ diatomic molecules	polar	non polar	reactive	non reactive	B
9	Sodium metal shows metallic luster is explained by	diffusion of Na ⁺	oscillation of loose electrons	excitation of free protons	strong crystal lattice	D
10	Which one of the characteristic of ionic solids?	high vapor pressure	good conductivity	low melting point	solubility in polar solvents	D
11	In closest packing of metal atoms, space occupied is	.67	.44	.74	.64	C
12	$\begin{array}{c} \delta+ & \delta- & \delta+ & \delta- \\ \text{H}-\text{Cl} & \cdots & \text{H}-\text{Cl} \end{array}$	Hydrogen bonding	dipole dipole forces	induced dipole forces	London forces	B
13	The type of hybridization in diamond is	spherical	sp ²	sp ³	sp-sp	C
14	A reaction with a tendency of occurring in forward and backward direction simultaneously is termed as	Irreversible	Unidirectional	Multidirectional	Reversible	D
15	The rate of reaction	Increases as the reaction proceeds	Decreases as the reaction proceeds	Remains unchanged	None of these	B
16	Rate of chemical reactions helps in designing industrial process which is	Completed instantaneously	Slow	Economical	None of these	C
17	Instantaneous rate of reaction is the rate at any	At Equilibrium	one instant	Given Temperature	Given Pressure	B
18	Exothermic Reactions heat is	Taken in	Give out to surroundings	Neither given nor lost	None of these	B
19	Anything under test or observation in laboratory is called	Surrounding	System	Confined space	None of these	B
20	Heat is the energy that flows across the	Surroundings	Boundaries of the system	only within the system	All of these	B
21	Conversion of Electrical energy into chemical energy is	Mechanical Energy	Electromechanical Energy	Electrochemistry	Chemical Energy	C
22	Decrease in Oxidation number is	Oxidation	Reduction	Both A and B	None of these	B
23	Where the attractive forces dominate the repulsive force, a state corresponding to the distance of _____ is reached.	89pm	79.4pm	75.4pm	8:00 PM	C
24	How much hydrogen atoms approach each other when potential energy decrease?	two	three	four	five	A
25	In periodic table, the ionization energies decrease from _____?	right to left	left to right	top to bottom	bottom to top	A
26	Group 1 elements are called?	alkaline earth metals	alkali metals	chalcogens	noble metals	B
27	Which of the following is non-metallic in nature	Nh	Tl	B	Ga	C
28	Which of the following is NOT a coinage element	Cu	Fe	Au	Ag	B
29	What are interstitial compounds?	Transition elements containing small proportion of H, B, C, N	Transition elements containing small proportion of He, Ne, Ar, Xe	The compounds in which transition elements are in their highest oxidation state.	The compounds of transition elements which are not stable	A
30	The formula of esters is	R - COOH	R - COR	R - COOR	R - COH	C
31	The number of isomers of hydrocarbon with thirty carbon atoms	7 billion	4 billion	2 billion	5 billion	B
32	Alicyclic organic compounds belong to	Heterocyclic compounds	Carbocyclic compounds	Saturated compounds	Aromatic compounds	B
33	Methane gas is also known as	Marsh gas	Pepper gas	Mustard gas	None of the above	A
34	If both double and triple bonds are present in the chain, double bond is given the _____	Higher number	Lower Number	Priority	Second position	B
35	Which of the following is NOT explained by Kekule structure of benzene	Hexagonal planar structure	The usual saturated behaviour of benzene	Unsaturation	Presence of three double bonds	B
36	In general formula of alkyl halide RX, X represents _____	Noble gases	Halogens	Metals	Metalloids	B
37	Which catalyst is used in the preparation of alkyl halide from alcohols	Zinc chloride	Pyridine	Sodium chloride	Thionyl chloride	A
38	CH ₃ CH ₂ CCl ₂ CH ₃ is called	3,3-Dichlorobutane	3,3-Trichlorobutane	2,2-Dichlorobutane	1,1-Dichlorobutane	C
39	The compounds which are formed by the replacement of one of the H of water by a alkyl group are called as _____?	Ethers	Phenols	Alcohols	Carboxylic acids	C
40	The alcohols in which the carbon which is attached to the OH group is further attached with two carbon atoms is called as?	Primary alcohols	Secondary alcohols	Tertiary alcohols	None of these	B
41	Which of the following alcohol does not convert into carboxylic acid when react with K ₂ Cr ₂ O ₇ /H ₂ SO ₄ ?	CH ₃ C(CH ₃) ₂ OH	CH ₃ CH(CH ₃)OH	CH ₃ OH	C ₂ H ₅ OH	A
42	Nucleophilic addition reactions of carbonyl compounds can be catalyzed by _____?	Acid	Base	Water	Both acid and base	D
43	Which of the following compound is present in camphor and menthone?	Aldehyde	Alcohol	Esters	Ketones	D
44	Cyanohydrins when reacted with mineral acid produces _____?	Beta Hydroxy carboxylic acids	Alpha hydroxy carboxylic acids	Carboxylic acids	Unsaturated acids	B
45	The compounds having -COOH group are called as _____?	Aldehydes	Ketones	Carboxylic acids	Alcohols	C
46	Which of the following is an example of Monocarboxylic acid?	Oxalic acid	Methanoic acid	Glutaric acid	Adipic acid	B
47	Which type of carboxylic acid is produced from the hydrolysis of nitriles?	Beta-hydroxy carboxylic acids	Acids having one carbon more than the starting material	Acids having one carbon less than the starting material	Alpha-hydroxy acids	B
48	Which of the following element is present in all proteins?	S	C	N	O	C
49	Which enzyme is used for diagnosis of Jaundice?	LDH-1	Protease	Alkaline phosphatase	None of these	C
50	Most of the enzyme reactions are _____?	Reversible	Irreversible	Condensation	Oxidation	A